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Interpret Scientific Illustrations Use Figure 5 and your knowledge of electromagnetic radiation to match the numbered items with the lettered items. The numbered items may be used more than once or not at

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Answers
all. a. longest
wavelength b.
highest frequency
c. greatest energy
1. gamma ray 2.
infraredwaves 3.
radio waves

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Although the speed

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of all electromagnetic waves in a vacuum is the same, waves can have different wavelengths and frequencies. As you can see from the equation on the previous page, wavelength and frequency are inversely related; in other words, as

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one quantity
increases, the
other decreases.

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the most valence
electrons for any
element is 8 (Noble
Gas Family). If an
atom has less than
that, it will try to

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gain, lose or share
valence electrons
with another
element in order to
have 8 valence
electrons.

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Chapter 5.1 to 5.3
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Models of the Atom
The Rutherford's
model of the atom
did not explain how

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an atom can emit
light or the
chemical
properties of an
atom.

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Section 5.2 –

Electron

Arrangement in

Atoms The electron
configuration of an
atom is the

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arrangement of the
electrons. There

are 3 rules that
govern the electron
configuration:

Aufbau's principle,
Pauli Exclusion
principle, and
Hund's rule.

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Electron
Configurations for
Elements in Period
Three Table 5-4
Figure 5-19. This
sublevel diagram
shows the order in
which the orbitals
are usually filled.
The proper
sequence for the
first seven orbitals
is 1s, 2s, 2p, 3s,
3p, 4s, and 3d.

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Study Guide. T/F

Like the visible spectrum, an atomic emission spectrum is a continuous range of colors.

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Figure 9 Chapter 5
electrons in atoms
answers 5.3. 1 Left:
a fragment of the
Tagish Lake
meteorite,
discovered in 2000
on the ice of Tagish
Lake, B. C. It is a
"stony" meteorite
that is dominated

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by ferromagnesian silicate minerals, and is similar in composition to Earth's mantle
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5.3.

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In Atoms Answers
5.3
Section 5.2
Quantum Theory

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and the Atom □

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□ Identify the relationships among a hydrogen atom's

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Models of the Atom
Rutherford used
existing ideas
about the atom

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and proposed an atomic model in which the electrons move around the nucleus, like the planets move around the sun.

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in Atoms . Name
Date 11. The
number of

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sublevels in an energy level is equal to the square of the principal quantum number of that energy level. 12. The maximum number of electrons that can occupy the fourth principal energy level of an atom is 32. 13. The higher the energy

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level occupied by
an electron the
more

cardinalnewman.en
school.org

This video
describes light as a
particle and wave.
It also describes
matter and
quantum of energy.

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Page 25/29

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in Atoms Pt 1

CHAPTER 5

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+ KEY Chemistry:
Matter and Change
1 Supplemental
Problems. 1.

Orange light has a
frequency of
 $4.8 \times 10^{14} \text{ s}^{-1}$.

What is the energy
of one quantum of
orange light? 2.

Which is greater,

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the energy of one photon of orange light or the energy of one quantum of radiation having a wavelength of 3.36×10^{-9} m? □3.

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Views. continued

their quest to understand atomic structure and the arrangement of electrons within atoms. Rutherford proposed that all of an atom's positive charge and virtually all of its

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mass are concentrated in a nucleus that is surrounded by fast-moving electrons ...

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