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A gas is a state of matter with no defined shape or

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volume. Gases have their own unique behavior depending on a variety of variables, such as temperature, pressure, and volume. While each gas is different, all gases act in a similar matter. This study guide highlights the concepts and laws

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dealing with the
chemistry of gases.

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Ten- Gases. Lecture
Notes 10.1
Characteristics of
Gases. •All
substances have
three phases: solid,

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liquid and gas.

- Substances that are liquids or solids under ordinary conditions may also exist as gases.
- These are often referred to as vapors.

Chapter Ten- Gases
#2 Pg 432 #5, 43, 45,
47, #3 Pg 432 #6 ...

Chemistry Gas

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kinetic theory. the particles in all forms of matter are always in motion. kinetic energy. energy of motion. solids. most of these are ionic compounds. ... unit cell. smallest group of particles within a crystal that retains the shape of the

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crystal. crystals.

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An Introduction to
Chemistry Section
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Section 13.1 Gases
and Their Properties
Goals To describe
the particle nature of

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both real and ideal gases. To describe the properties of gases that can be used to explain their characteristics: volume, number of particles, temperature, and pressure.

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CP Chemistry Unit 5
(gas laws) identify
the variables that
impact gas.... identify
the relationship
between each....
explain at the
particle level the
relat.... apply gas
laws to novel
situations. pressure,
number, volume, tem
perature. all others
are linear.

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Gas Laws STUDY

GUIDE Due:

February 12th Units

of Measurement: For

the following

questions, use the

following answer

choices to indicate

what each unit of

measurement is

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used to measure. A.

Pressure B. Volume

C 1. K 4. kPa A 2.

atm 5. L 3. mL 6.

°C C. Temperature

A 7. A 8.

Gas Laws STUDY

GUIDE Due:

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Unit 1: Matter Video

1 This video will

introduce you to the

general properties

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of Solids, Liquids,
Gases and will teach
you how to
effectively draw
their particle
models.

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Properties & Particle
Models

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Study Guide:
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Write. Spell. Test.

PLAY. Match.

Gravity. Created by.

Ads827. Key

Concepts: Terms in
this set (10)

sublimation. A

change directly from

the solid to the

gaseous state

without becoming

liquid. deposition.

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the change of state
from a gas directly
to a solid without ...

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Start studying
Chemistry B Unit 1:
Solids, Liquids, and
Gases. Learn
vocabulary, terms,
and more with

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and other study
tools.

Chemistry B Unit 1:
Solids, Liquids, and
Gases Flashcards ...
Know units of
pressure in atm,
torr, mm Hg, and
kPa, Calculate using
Boyle's Law,
Charles' Law, Gay-
Lussac's Law,

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Answers
Combined Gas Law,
and Ideal Gas Law
(using 0.0821 for R).
Change gases to
STP. Calculate
partial pressure of a
gas using Dalton's
Law.

Study Guide for AP
Chemistry – Chapter
5, Gas Laws
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Notes 101

Characteristics of
Gases •All

substances have
three phases: solid,
liquid and gas

- Substances that
are liquids or solids
under ordinary
conditions may also
exist as gases

- These are often
referred to as

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vapors • Many of the properties of gases differ from those of solids

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law of combining
volumes - at given p
ressure/temperature
, volumes of
reacting gases exist
in simple ratios

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equal volumes of
gases at same temp
erature/pressure
have equal numbers
of molecules volume
of gas (at constant p
ressure/temperature
) directly
proportional to
number of moles of
gas $V = \text{constant} \times n$

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expressed by the
formula $PV = nRT$
where P = pressure
 V = volume n =
number of moles of
gas R = ideal gas
constant T =
absolute
temperature The
value of R depends
on the units of
pressure, volume
and temperature.

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kinetic molecular

model, gas particles

are in constant

random motion,

straight line motion,

they are of

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negligible volume,
they have no forces
of attraction, and
they have elastic
collisions.

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model, gas particles
are in constant
random motion,
straight line motion,
they are of
negligible volume,
they have no forces
of attraction, and
they have elastic
collisions.

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